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VERIFICATION OF TRANSLATION

Sir:

I, Isamu Ogoshi, having been warned that willful false statements and the like are punishable by fine or imprisonment or both, under section 1001 of Title 18 of the United States Code, and may jeopardize the validity of the above-captioned application and any patent issuing thereon, declare:

(1) I am a patent attorney authorized to practice law in Japan and am engaged in the practice of law with OGOSHI International Patent Office at Toranomom 9 Mori Bldg. 3F, 2-2, Atago 1-Chome, Minato-ku, Tokyo 105-0002, Japan.

(2) I am fluent in the Japanese and English Languages.

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- (3) I have reviewed the attached translation, and certify that it is an accurate

English translation of the Japanese language international application of Yuichiro Shindo and Kouichi Takemoto filed on July 28, 2004 and given International Application No. PCT/JP2004/010714.

- (4) All of the statements made herein of my own knowledge are true and all statements made herein on information and belief are believed to be true.

January 30, 2006
Date

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HIGH PURITY COPPER SULFATE
AND METHOD FOR PRODUCTION THEREOF

TECHNICAL FIELD

The present invention relates to a manufacturing method of high purity copper sulfate including the steps of dissolving commercially available copper sulfate crystals (purity, for instance, is 95 to 99.9wt%) in purified water or acid, and thereafter removing impurities, and to the high purity copper sulfate obtained thereby. The starting raw material does not have to be copper sulfate crystals, and may be a resultant solution of dissolving copper in acid containing sulfuric acid, or copper sulfate crystals manufactured therefrom.

BACKGROUND ART

Copper sulfate (Cu_2SO_4) is a white-colored powder, but this generally refers to as a pentahydrate ($\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$), and is an azurite blue crystal.

Copper sulfate is being used as an electrolytic solution, pigment, insecticide, antiseptic, mordant, battery material, medicine and so on. In particular, when copper sulfate is to be used as the electroplating solution for electronic components such as a semiconductor device, high purity copper sulfate is sought after. Commercially available copper sulfate has a purity level of 95 to 99.9wt%, and it is necessary to purify this further in order to realize a purity level of 4N to 5N or higher.

As conventional technology, disclosed is a method of obtaining copper

sulfate with low Ni content by using electrolytic copper powder, which was electrodeposited and collected from an electrolytic solution, as the raw material, dipping this in acid to selectively dissolve and remove Ni, filtering the resultant solution and thereafter dissolving the copper powder in sulfuric acid to realize crystallization (e.g., refer to Japanese Patent Laid-Open Publication No. 2001-10817).

Further, also disclosed is technology for obtaining copper sulfate with a low nickel content by making copper sulfate containing nickel into an aqueous solution, heating this to 80°C or higher, and collecting the copper sulfate crystals that are separated and settled as a result of the heating in order to realize concentrated recrystallization (e.g., refer to Japanese Patent Laid-Open Publication No. 2001-31419).

DISCLOSURE OF THE INVENTION

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An object of the present invention is to provide a manufacturing method of high purity copper sulfate capable of efficiently removing impurities at a low cost by dissolving commercially available copper sulfate crystals in purified water or acid and thereafter subjecting this to the refining process, and high purity copper sulfate obtained thereby.

The present invention provides:

- 1) High purity copper sulfate wherein the content of Ag impurities is 1wtppm or less, and having a purity of 99.99wt% or higher;
- 2) High purity copper sulfate wherein the metalloid element impurities of As, Sb and Bi are respectively 1wtppm or less, and having a purity of 99.99wt% or higher;
- 3) High purity copper sulfate according to paragraph 1) above, wherein the metalloid element impurities of As, Sb and Bi are respectively 1wtppm or

less, and having a purity of 99.99wt% or higher;

4) High purity copper sulfate wherein undissolved residue of $1\mu\text{m}$ or greater exists in an amount of 100 residues/L or less, and having a purity of 99.99wt% or higher;

5) High purity copper sulfate according to any one of paragraphs 1) to 3) above, wherein undissolved residue of $1\mu\text{m}$ or greater in water exists in an amount of 100 residues/L or less, and having a purity of 99.99wt% or higher;

6) High purity copper sulfate according to any one of paragraphs 1) to 5) above, having a purity of 99.999wt% or higher;

7) A manufacturing method of high purity copper sulfate including the steps of dissolving crude copper sulfate crystals or copper metal, and subjecting this to solvent extraction and thereafter to active carbon treatment in order to realize recrystallization; and

8) A manufacturing method of high purity copper sulfate according to any one of paragraphs 1) to 6) above, including the steps of dissolving crude copper sulfate crystals or copper metal, and subjecting this to solvent extraction and thereafter to active carbon treatment in order to realize recrystallization.

Effect of the Invention

The present invention yields a superior effect in that it is capable of efficiently removing impurities by dissolving commercially available copper sulfate having a purity level of 95 to 99.9wt% in purified water or acid and thereafter subjecting this to solvent extraction and active carbon treatment, and thereby manufacturing high purity copper sulfate at a low cost.

BRIEF DESCRIPTION OF THE DRAWINGS

FIG. 1 is a diagram showing the flow of the manufacturing method of high purity copper sulfate.

BEST MODE FOR CARRYING OUT THE INVENTION

Examples of analytical values of commercially available copper sulfate are shown in Table 1. As shown in Table 1, copper sulfate contains transition metals such as iron, nickel and cobalt, and impurities such as Ca, Cr, Al and Si in amounts of roughly 0.5 to 100wtppm, respectively. Further, roughly several ppm of Ag, As, Sb and Bi are also contained therein.

In addition to the above, foreign matters such as impurities or oxides; for instance, several to several ten wtppm of alkali elements such as Na and K; and several ten wtppb of radio active elements such as U and Th are also contained therein.

Although it is also possible to dissolve commercially available copper metal or copper scrap in sulfuric acid and to use this as copper sulfate, in such a case, the foregoing impurities will still be contained therein.

As described above, particularly when using copper sulfate as the electroplating solution for electronic components such as a semiconductor device, it is necessary to use copper sulfate having a high purity of 99.99wt% or higher. In particular, since Ag will become electrodeposited on the plating film and cause problems such as contamination, it must be reduced.

Ag is very similar to Cu, and is an element that is difficult to remove. Further, since it does not have any adverse effect on the electrical conduction property, conventionally, this was never intentionally removed. Nevertheless, a defect has been discovered in that the slight content of Ag in

the electroplated film would cause the plated film to become hard and brittle.

With the plating wiring having this kind of defect, problems would arise in that such wiring would break or disconnect under stress. Therefore, in order to avoid this kind of defect, it is necessary to reduce Ag as much as possible; that is, to reduce Ag to 1wtppm or less.

Further, metalloid element impurities of As, Sb and Bi are easily engulfed in the plated film, and must be reduced since they similarly cause the contamination of the plated film.

When As, Sb and Bi are engulfed in the thin film wiring, there is a problem in that the electrical conduction property will become inferior, and the resistivity will become high. Therefore, it is desirable to make these impurities to be within the foregoing scope.

In addition to the above, impurities such as organic matter and foreign matter are also contained, and these will become undissolved residue in water, and similarly cause the contamination of the plated film. Thus, these impurities must be reduced to 100 residues/L or less. Specifically, this undissolved residue caused problems the thinner the wiring width of the semiconductor became. This was overlooked in the past.

In particular, upon embedding in a trench, there was a problem in that the embedding could not be conducted or would be insufficient as a result of being obstructed by foreign matter. In order to reduce the generation of defective goods, the existence of undissolved residue is a problem that can no longer be overlooked, and it is therefore desirable to make such undissolved residue to be within the foregoing scope.

Generally speaking, it is desirable that the conditions for reducing Ag, reducing As, Sb, Bi and undissolved residue are all fulfilled, but there are cases that the conditions would not cause a particular problem depending on the purpose or characteristic of the wiring. In other words, such conditions

may be suitably selected depending on which characteristic is important. Therefore, the conditions may be adjusted independently. The present invention covers all such cases.

5 The foregoing copper sulfate is dissolved in purified water or acid such as diluted sulfuric acid at room temperature, and this is thereafter subject to solvent extraction. Further, active carbon treatment is performed to remove impurities such as organic matter and foreign matter as well as noble metals such as Ag. The removal of impurities such as organic matter and foreign matter and oil films, and the removal of noble metals such as Ag via active
10 carbon treatment do not exist in conventional technology, and such removal process is a new discovery and method provided by the present invention.

Thereafter, impurities such as organic matter and foreign matter are filtered and removed. When there is undissolved foreign matter and the like, these are also filtered and removed.

15 Next, the solution or filtrate that dissolved copper sulfate is heated to a temperature of roughly 50 to 100°C, and subject to evaporative concentration. After performing evaporative concentration, this is cooled to room temperature, and crystals are deposited thereby.

20

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[Table 1]

wtppm					
Impurity Element	Raw Material	Example 1	Example 2	Example 3	Comparative Example 1
B	1	<0.01	<0.01	<0.01	0.7
Na	13	<0.1	0.7	<0.1	7
Mg	0.8	<0.01	<0.01	<0.01	0.5
Al	3.2	<0.1	<0.1	<0.1	1.6
Si	17	0.5	0.7	0.2	8
P	2.3	0.3	<0.1	0.1	1.1
Cl	8.5	0.5	0.7	<0.1	3.5
K	3.1	<0.1	0.3	<0.1	1.2
Ca	3.5	<0.1	0.8	<0.1	1.0
Ti	0.4	<0.01	<0.01	<0.01	0.3
Cr	18.1	<0.1	0.9	<0.1	9.5
Mn	0.2	<0.1	<0.1	<0.1	0.1
Fe	81	<0.1	7.6	<0.1	8.5
Co	3.5	<0.1	0.35	<0.1	1.5
Ni	4.3	<0.1	0.43	<0.1	1.2
Ag	2.5	<0.01	0.1	<0.01	1.5
As	4.8	<0.1	0.32	<0.1	3.2
Sb	2.4	<0.1	0.15	<0.1	1.3
Bi	2.1	<0.1	0.18	<0.1	1.8
Th	0.02	<0.001	0.001	<0.001	0.01
U	0.01	<0.001	0.001	<0.001	0.007
Undissolved Residue in Water (Residue /L)	150	None	None	None	110

Thereafter, this is cooled to room temperature, crystals of copper sulfate are deposited, and these are filtered to obtain blue-colored high purity copper sulfate. The initial crystals may be removed as necessary.

It is desirable that the filtrate after the ultimate filtration is 2 to 40wt% of the original fluid volume. This is in order to prevent impurities such as Na and K from getting mixed into the crystals by remaining non-crystallized copper sulfate in a residual liquid.

The drying temperature is preferably 40 to 100°C. If the temperature is less than 40°C, too much time will be required for removing the adsorption moisture, and if the temperature exceeds 100°C, this is not preferable since the adhesive moisture of pentahydrates will be removed and the form of copper sulfate will change. The impurities after the foregoing refining process are similarly shown in Table 1.

The various impurities shown in Table 1 are reduced respectively to 1ppm or 0.1ppm or less, and high purity copper sulfate having a purity level of 4N to 5N is obtained. Incidentally, among the elements shown in Table 1, although there are certain elements (for instance Na, Si, Cl, Ca, Cr, Fe, Co and Ni) that are contained in amounts slightly more than Ag or As, Sb and Bi, these impurities will remain in the bath as ions when using these impurities as a copper sulfate solution, and it is unlikely that these will enter the film wiring. Thus, no problems will arise with the level of these impurities shown in Table 1. The flow of the manufacturing method of high purity copper sulfate according to the present invention is shown in FIG. 1.

Examples

The Examples are now explained. Incidentally, these Examples are merely illustrative, and the present invention shall in no way be limited thereby. In other words, the present invention shall only be limited by the

scope of claim for a patent, and shall include the various modifications other than the Examples of this invention.

(Example 1)

5 250g of commercially available copper sulfate crystals ($\text{Cu}_2\text{SO}_4 \cdot 5\text{H}_2\text{O}$) having a purity level of 99.9wt% and containing the impurities shown in Table 1 were dissolved in 1000ml of diluted sulfuric acid at room temperature.

10 Next, this was subject to solvent extraction with D2EHPA. This was further subject to active carbon treatment in order to remove impurities such as organic matter and Ag as well as foreign matter. This was thereafter heated to 900°C or higher, and water was evaporated in a prescribed amount. Crystals started to appear from around 700ml, and ended at 200ml.

15 Next, this was cooled to room temperature in order to deposit refined copper sulfate crystals, and the crystals were filtered. Since there is a risk that impurities such as Na and K could get mixed into the refined copper sulfate when evaporation is increased extremely, evaporation was stopped midway in order to obtain a residual liquid.

20 When this high purity copper sulfate was redissolved and was filtered, there was no undissolved residue of 1 μm or greater in water. The results are shown in Table 1.

25 As shown in Table 1, the various impurities listed in Table 1 have been reduced respectively to 1wtppm or 0.1wtppm or less, and high purity copper sulfate having a purity level of 4N to 5N or higher was obtained. With the processes shown in Example 1 of the present invention, a significant improvement in purity has been confirmed.

 These impurities are particularly disfavored in the copper plating to circuits or wirings upon manufacturing a semiconductor device or the like, and the reduction of such impurities is extremely effective.

(Example 2)

The same raw material as with Example 1 was used, and, without
subjecting this to solvent extraction, only the active carbon treatment and the
subsequent processes of Example 1 were performed. The results are
5 shown in Table 1.

As shown in Table 1, although slightly inferior to Example 1, only Fe
was 7.6wtppm, and the other impurities have been reduced to 1wtppm or
0.1wtppm or less, and high purity copper sulfate having a purity level of 4N to
5N or higher was obtained. Further, when this high purity copper sulfate
10 was redissolved and was filtered, there was no undissolved residue of 1 μ m or
greater in water.

With the processes shown in Example 2 of the present invention, a
significant improvement in purity has been confirmed. These impurities are
particularly disfavored in the copper plating to circuits or wirings upon
15 manufacturing a semiconductor device or the like, and the reduction of such
impurities is extremely effective.

(Example 3)

Copper scrap raw material having the same purity level as with
Example 1 was used, and the same processes as with Example 1 were
20 performed. The results are shown in Table 1.

As shown in Table 1, impurities have been reduced to 1wtppm or
0.1wtppm or less, and high purity copper sulfate having a purity level of 4N to
5N or higher was obtained. Further, when this high purity copper sulfate
was redissolved and was filtered, there was no undissolved residue of 1 μ m or
25 greater in water when.

With the processes shown in Example 3 of the present invention, a
significant improvement in purity has been confirmed. These impurities are
particularly disfavored in the copper plating to circuits or wirings upon

manufacturing a semiconductor device or the like, and the reduction of such impurities is extremely effective.

(Comparative Example 1)

The same raw material as with Example 1 was used, and, without
5 subjecting this to solvent extraction and active carbon treatment, the same processes as with Example 1 were performed for recrystallization. The results are shown in Table 1.

As shown in Table 1, reduction of impurities was hardly achieved. Further, when this high purity copper sulfate was redissolved and was filtered,
10 undissolved residue of $1\mu\text{m}$ or greater in water existed in an amount of 110 residues/L.

The improvement in purity could not be achieved with the Comparative Example, and these impurities caused problems in that impurities which are particularly disfavored in the copper plating to circuits or wirings upon
15 manufacturing a semiconductor device or the like remained in large quantities.

Industrial Applicability

The present invention yields a superior effect in that it is capable of
20 efficiently removing impurities by dissolving commercially available copper sulfate having a purity level 95 to 99.9wt% in purified water or acid and thereafter subjecting this to solvent extraction or active carbon treatment, and thereby manufacturing high purity copper sulfate at a low cost, and is extremely useful as an electrolytic solution, pigment, insecticide, antiseptic,
25 mordant, battery material, medicine and so on, as well as the electroplating solution for electronic components such as a semiconductor device.